

Figure 5. When the rate equation is considered as second order in acetaldehyde concentration, the slopes of the lines on this plot decrease with increasing acetaldehyde concentration instead of remaining constant as a true second-order case would require

## **ACKNOWLEDGMENT**

The authors are grateful for financial support of this project by the National Science Foundation and the U.S. Industrial Chemicals Corporation.

## LITERATURE CITED

- (1) Bell, R.P., J. Chem. Soc. 1937, p. 1637.
- (2) Bell, R.P., Trans. Faraday Soc. 37, 707 (1941).
- (3) Bell, R.P., Clurrie, J.C., Ibid., 48, 439 (1952).
- (4) Benton, C.H., Ph.D. thesis in chemistry, University of Illinois, Urbana, Ill., 1951.
- Bonhoeffer, K.F., Walters, W.D., Z. physik. Chem. A-181, 441 (1938).
- (6) Hann, A., Lapworth, A., J. Chem. Soc. 85, 46 (1904).
- (7) Herbert, J.B.M., Lauder, I., Trans. Faraday Soc. 34, 432 (1938).
- (8) Matsuyama, H., Proc. Japan. Acad. 27, 552 (1951).
- (9) Siggia, S., Maxcy, W., Anal. Chem. 19, 1023 (1947).
- (10) Usherwood, E.H., J. Chem. Soc. 123, 1717 (1923).

RECEIVED for review November 2, 1959. Accepted March 18, 1960.

## CORRECTION

In "Phase Equilibria in Hydrocarbon Systems. Volumetric and Phase Behavior of Ethane-n-Pentane System"

[H.H. Reamer, B.H. Sage, and W.N. Lacey, J. Chem. Eng. Data 5, 44 (1960)] these figures should have appeared instead of the ones on page 49.

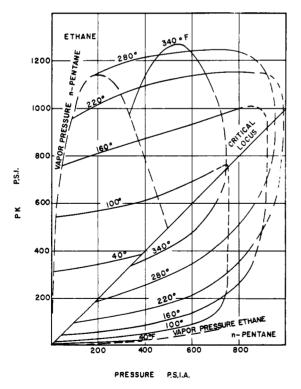


Figure 3. Equilibrium ratios for ethane and n-pentane

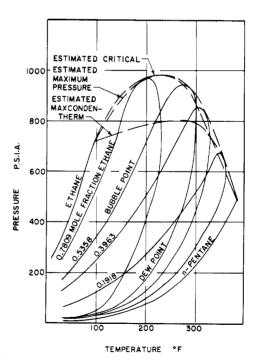


Figure 4. Pressure-temperature diagram for the ethane—n-pentane system